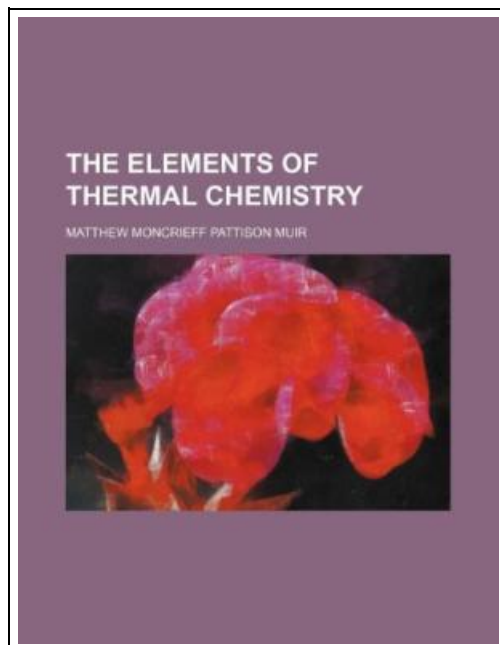


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THE ELEMENTS OF THERMAL CHEMISTRY



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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1885 Excerpt: .sketch indicates that there is a more or less well-marked connection between the phenomena of evaporation and dissociation. This connection will be discussed later on. (See pars. 136, and 145--152.) 129. Melting. In order to change a given mass of a solid at a specified temperature into the same mass of liquid at the same temperature, pressure being constant throughout the change, a definite quantity of heat must be added to the body. This quantity of heat is called the heat of liquefaction, or heat of fusion, or latent heat, of the solid. The value of the heat of liquefaction of a body may be determined by bringing equal masses of the body, at definite temperatures, one above and one below the melting point, into a calorimeter containing a definite amount of some liquid which dissolves the given solid, and measuring the quantities of heat produced during each process of solution. As in each case the final state of the system is the same, and as the only difference between the initial states is that one system is solid and the other liquid, the difference between the quantities of heat produced (or absorbed) in the process measures the quantity of heat which must be added to the initial solid system to cause it to pass into the initial liquid system. The following data will serve as examples of the application of this method: --Hydrated sulphuric acid. Heat of solution in water. solid H₂SO₄, 400H₂O = 7,120;...



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