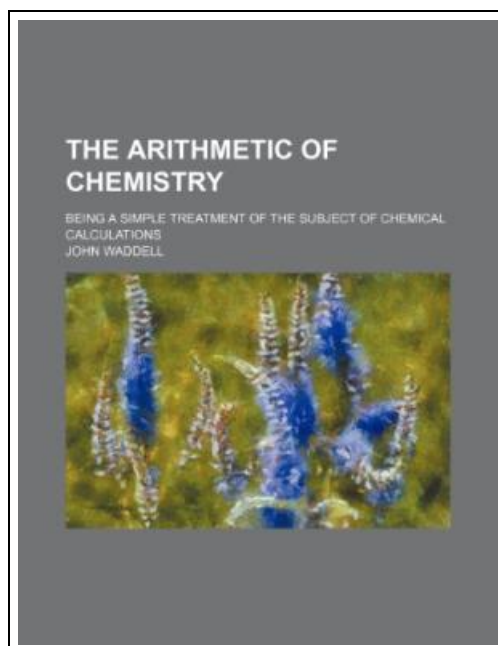


The Arithmetic of Chemistry; Being a Simple Treatment of the Subject of Chemical Calculations



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Reviews

This book is really gripping and fascinating. Of course, it is actually play, nonetheless an interesting and amazing literature. You will not feel monotony at anytime of the time (that's what catalogs are for about if you request me).
(Delbert Gleason)

THE ARITHMETIC OF CHEMISTRY; BEING A SIMPLE TREATMENT OF THE SUBJECT OF CHEMICAL CALCULATIONS



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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1915 Excerpt: .1898.) (0) Density is compared to air, therefore the formula weight-- $28.8 \times 1.524 = 43.9$. /i 28 W 8 = 0.973 (f) $1.97 \times 22.4 = 44.1$. 18. What volume of chlorine at normal temperature and pressure would be required (a) to combine with 10 litres of olefiant gas; (6) to decompose 10 litres of sulphuretted hydrogen liberating the sulphur; (c) to decompose 10 grammes of potassium iodide? (London Matric., June, 1889.) (a) 10 litres; (6) 10 litres; (f) 0.675 19. How many litres of oxygen, measured at 150 C. and 750 mm. pressure, are required for the complete combustion of 100 grammes of marsh gas? (Edin. Univ., 1st Professional, 1890.) Ans. 299.3 litres. 20. What volume in litres, of oxygen measured at 730 mm. and 10 C, can be obtained from 1 kilogramme of pyrolusite, containing 80 of peroxide of manganese? (Edin. Univ., 1st Professional, 1873.) Ans. 74.1 litres. 21. What volume of sulphurous acid gas, measured at 740 mm. pressure and 150 C. temperature, can be obtained by burning 10 grammes of sulphur? (Edin. Univ., 1872.) Ans. 7.57 litres. 22. (a) Find the density of the gas whose molecular formula is C_2N_2 . (6) Find the weight of a litre of hydrochloric acid. (c) 1 litre of a gas weighs 1.25 grammes under standard conditions; find its molecular weight. (Cornell Univ., September, 1896.) Ans. (-) 1.807; () 1.629 grammes; (0 27.9-23. How many litres of dry atmospheric air at 740 mm. pressure and 15 C....



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