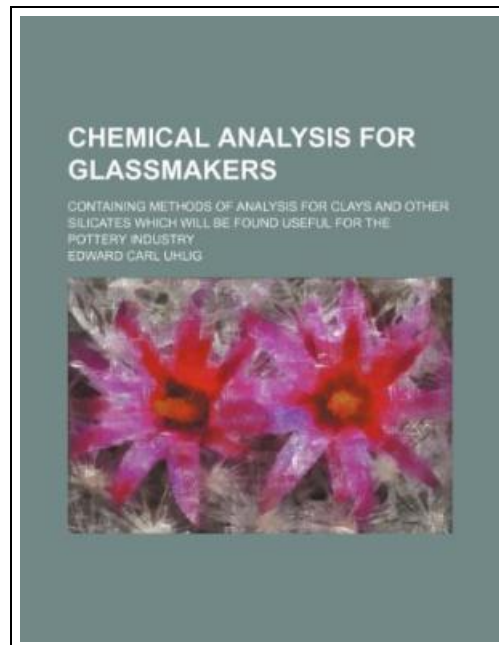


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CHEMICAL ANALYSIS FOR GLASSMAKERS; CONTAINING METHODS OF ANALYSIS FOR CLAYS AND OTHER SILICATES WHICH WILL BE FOUND USEFUL FOR THE POTTERY INDUSTRY



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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1903 Excerpt: . r. When the effervescence has subsided, close the stop-cock in / and place the lighted burner d under a. Permit the flame to just touch the bottom of a, and gradually heat to boiling. Allow to boil for about-two minutes, and extinguish the flame. Open the stop-cock in / and partly close v, allowing the water to drop rapidly. It will be noticed that the part of the absorption apparatus r which contains the Soda Lime has become quite warm. Continue drawing a current of air through until this last-mentioned tube does not feel warm when touched by the hand. Close v, remove r and insert the caps. Hang it on the balance hook and weigh it. The increase in weight will be the Carbon Dioxide (CO₂) originally contained in the Bicarbonate. Determination of Alkali, Na₂O: Weigh 3.1 grams of the Bicarbonate and brush it into a 100 Cc. Erlenmeyer flask. Add about 20 Cc. of cold water and a drop of Methyl Orange Indicator. Remove to a burette and titrate with Standard Acid, as described under Soda Ash. The number of Cc. and tenths used of the Standard Acid will represent the percentage of Alkali (Na₂O) present. Calculation. Having determined the constituents and impurities of the Bicarbonate, the next step is to calculate the percentage of actual Sodium Bicarbonate which is available. The method of calculation is best shown by an example. Example of Sodium Bicarbonate Analysis: Determination of Moisture, 10 grams taken: Weight...

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